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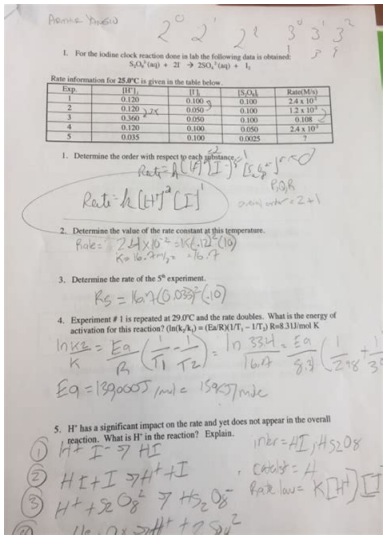
so many fake sites. this is the first one which worked! Many thanks



1. For the iodine clock reaction done in lab the following data is obtained:
$S_2O_8^{2-}(aq) + 2I^- \rightarrow 2SO_4^{2-}(aq) + I_2$

Rate information for 25.0°C is given in the table below.

Exp.	$[H^+]$	$[I^-]$	$[S_2O_8^{2-}]$	Rate(M/s)
1	0.120	0.100	0.100	2.4×10^{-3}
2	0.120	0.050	0.100	1.2×10^{-3}
3	0.360	0.050	0.100	0.108
4	0.120	0.100	0.050	2.4×10^{-3}
5	0.035	0.100	0.0025	?

1. Determine the order with respect to each substance.
2. Determine the value of the rate constant at this temperature.
3. Determine the rate of the 5th experiment.
4. Experiment # 1 is repeated at 29.0°C and the rate doubles. What is the energy of activation for this reaction? $(\ln(k_2/k_1) = (Ea/R)(1/T_1 - 1/T_2)$ R=8.31J/mol K
5. H^+ has a significant impact on the rate and yet does not appear in the overall reaction. What is H^+ in the reaction? Explain.

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The Rate Of An Iodine Clock Reaction Lab Answers